GOVT. DEGREE COLLEGE(MEN) SRIKAKULAM

AFFIDIVT TO DR.B.R.AMBETKAR UNIVERSITY

CERTIFICATED

DEPARTMENT OF CHEMICAL SCIENCE

PROJECT ON

MAKING POTASH ALUM

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DEDICATION :

MY FATHER
MR R.NARISIMHA MURTHY
&
MY MOTHER
MRS. ANURADHA

INTRODUCTION:

Aluminium because of its low density, high tensile strength and resistance to corrosion is widely used for the manufacture of aeroplanes, automobiles lawn furniture as well as for aluminium cans. Being good conductor of electricity it is used for transmission of electricity. Aluminium is also used for making utensils. The recycling of aluminium cans and other aluminium products is a very positive contribution to saving our natural resources. Most of the recycled aluminium is melted and recast into other aluminium metal products or used in the production of various aluminium compounds, the most common of which are the alums. Alums are double sulphates having general formula

$X_2SO_4.M_2(SO_4)_3.24H_2O$

X = Monovalent cation; M = Trivalent cation

Some important alum and their names are given below:

$K_2SO_4.Al_2(SO_4)_3.24H_2O$ – Potash Alum

$Na_2SO_4.Al_2(SO_4)_3.24H_2O$ – Soda Alum

Introduction

$K_2SO_4.Cr_2(SO_4)_3.24H_2O$ – Chrome Alum

$(NH)_2SO_4.Fe_2(SO_4)_3.24H_2O$ – Ferric Alum

Potash alum is used in papermaking, in fire extinguishers, in food stuffs and in purification of water soda alum used in baking powders and chrome alum is used in tanning leather and water proofing fabric.
In addition to these primary uses, alum is also used as

1. *An astringent* a substance or preparation that draws together or constricts body tissues and is effective in

stopping the flow of blood or other secretions. Alum has also been used by conventional hairdressers for treating shaving cuts,

1. *A mordant* substances used in dyeing to fix certain dyes on cloth. Either the mordant (if it is colloidal) or a colloid produced by the mordant adheres to the fiber, attracting and fixing the colloidal mordant dye. The insoluble, colored precipitate that is formed is called a lake. Alum is a basic mordant used for fixing acid dyes.

INTRODUCTION:

1. *For the removal of phosphate from natural and waste waters* the aluminium ions of alum combine with the orthophosphate around pH 6 to form the solid aluminum hydroxyphosphate which is precipitated and
2. *For fireproofing fabrics* The major uses of alums are based on two important properties, namely precipitation of $Al(OH)_3$ and those related to the acidity created by the production of hydrogen ions.

$$Al(H_2O)_6^{+3} \rightarrow Al(OH)_3 \downarrow + 3H_2O + 3H^+$$

The H^+ ions generated are used foe reacting with sodium bicarbonate to release CO_2. This property is made use of in baking powder and CO_2 fire extinguishers.

AIM :

To prepare potash alum from aluminium scrap

REQUIREMENT:

1. 250 ml flask
2. Funnel
3. Beaker
4. Scrap aluminium or cola can
5. Potassium hydroxide solution (KOH)
6. 6 M Sulphuric Acid (H_2SO_4)
7. Water Bath
8. Ethanol

THEORY:

Aluminum metal is treated with hot aqueous KOH solution. Aluminium dissolves as potassium aluminate, $KAl(OH)_4$. salt.

$2Al_{(s)} + 2KOH_{(aq)} + 6H_2O_{(l)}$ ¾® $2KAl(OH)_{4\,(aq)} + 3H_2$

Potassium aluminate solution on treatment with dil. Sulphuric acid first gives precipitate $Al(OH)_3$, which dissolves on addition of small excess of H_2SO_4 and heating.

$2KOH_{(aq)} + H_2SO_{4(aq)}$ ¾® $2Al(OH)_{3\,(s)} + K_2SO_{4(aq)} + 2H_2O_{(l)}$

$2Al(OH)_{3\,(s)} + 3\ H_2SO_{4(aq)}$ ¾® $Al_2(SO_4)_{3(aq)} + 6H_2O_{(l)}$

The resulting solution is concentrated to near saturation and cooled. On cooling crystals of potash alum crystallize out.

$K_2SO_{4(aq)} + Al_2(SO_4)_{3(aq)} + 24H_2O_{(l)}$ ¾® $K_2SO_4.Al_2(SO_4)_3.\ 24H_2O_{(s)}$

REACTIONS :

$2Al_{(s)} + 2KOH_{\ (aq)} + 6H_2O_{\ (l)}$ ¾® $2KAl\ (OH)_{4\ (aq)} + 3H_2$

$K_2SO_{4(aq)} + Al_2(SO_4)_{3(aq)} + 24H_2O_{(l)}$ ¾® $K_2SO_4.Al_2(SO_4)_3.$

$24H_2O_{(s)}$

$2Al(OH)_{3\,(s)} + 3\ H_2SO_{4(aq)}$ ¾® $Al_2(SO_4)_{3(aq)} + 6H_2O_{(l)}$

$2KOH_{(aq)} + H_2SO_{4(aq)}$ ¾® $2Al(OH)_{3\,(s)} + K_2SO_{4(aq)} + 2H_2O_{(l)}$ $K_2SO_{4(aq)} +$
$Al_2(SO_4)_{3(aq)} + 24H_2O_{(l)}$ ¾® $K_2SO_4.Al_2(SO_4)_3.\ 24H_2O_{(s)}$

PROCEDURE :

- Clean a small piece of scrap aluminium with steel wool and cut it into very small pieces. Aluminium foil may be taken instead of scrap aluminium.
- Put the small pieces of scrap aluminium or aluminium foil (about 1.00g) into a conical flask and add about 50 ml of 4 M KOH solution to dissolve the aluminium.
- The flask may be heated gently in order to facilitate dissolution. Since during this step hydrogen gas is evolved this step must be done in a well ventilated area.
- Continue heating until all of the aluminium reacts.
- Filter the solution to remove any insoluble impurities and reduce the volume to about 25 ml by heating.
- Allow the filtrate to cool. Now add slowly 6 M H_2SO_4 until insoluble $Al(OH)_3$ just forms in the solution. Procedure
- Gently heat the mixture until the $Al(OH)_3$ precipitate dissolves.
- Cool the resulting solution in an ice-bath for about 30 minutes whereby alum crystals should separate out. For better results the solution may be left overnight for crystallization to continue.
- In case crystals do not form the solution may be further concentrated and cooled again.
- Filter the crystals from the solution using vacuum pump, wash the crystals with 50/50 ethanol-water mixture.
- Continue applying the vacuum until the crystals appear dry.
- Determine the mass of alum crystals.

OBSERVATION

Observation

Mass of aluminium metal	=	1 g
Mass of potash alum	=	12.92g
Theoretical yield of potash alum	=	17.56g
Percent yield	=	73.49%

RESULT:

Potash alum of 73.49% yield was prepared from aluminium scrap.